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2. atoms of the same element are identical; atoms of different elements differ in size, mass, and other properties 3. atoms cannot be subdivided, created, or destroyed 4. atoms of different elements combine in simple whole-number ratios to form chemical compounds 5. atoms are only combined, separated, or destroyed in chemical

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reactions
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Chapter Test Name
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Test A, continued

_____13. To calculate
the number of atoms
present in 2.0 mol of
an element, you would
a. add Avogadro's
number of atoms per
mole to 2.0 mole. b.
subtract Avogadro's

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number of atoms per
mole from 2.0 mole. c.
multiply Avogadro's
number of atoms per
mole by 2.0 mole.

**Assessment Chapter
Test A - Kettering
City School District**

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Chapter Test Name

Class Date Chapter

Test B, continued Write

the orbital notation for
the following elements
in the space provided.

37. lithium, atomic

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number 3 38. carbon,
atomic number 6 39.
neon, atomic number
10 PART VI Write the
answers to the
questions on the line to
the left, and show your
work in the space
provided.

**Assessment Chapter
Test B - Ed W. Clark
High School**

Particle Mass number
19. Proton 20. Neutron
21. Electron 10

CHAPTER 3 TEST
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Class CHAPTER 3 TEST
continued SHORT
ANSWER Write the
answers to the
following questions in
the space provided. 22.

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Chemistry Chapter3
Practice Test**

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CHAPTER 3 TEST
Atoms: The Building
Blocks of Matter ... 10
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**Chapter 4 Modern
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**Arrangement Of
Electrons In ...**

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Chapter Test Name

Class Date Chapter

Test A, continued _____

4. The electron configuration below violates a. the Pauli exclusion principle. b. the Aufbau principle. c. Hund's rule. d. Both (a) and (c) _____

5. A

photon is emitted from a gaseous atom when an electron moves to its ground state from

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a(n) a. inner shell. b ...

Test Answers

**Assessment Chapter
Test A -**

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CHAPTER 4 REVIEW

Arrangement of
Electrons in Atoms
SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. In what way does the photoelectric effect support the particle theory of light?

**4 Arrangement of
Electrons in Atoms**

CHAPTER 3 TEST Class

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Chapter Atoms
The Atom

Atoms: The Building Blocks of Matter
MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice best completes the statement or answers the question. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of

a. energy

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Mixed Questions. CH 1
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Chapter Three Atoms:
Building Blocks of
Matter. Objectives:
Chapter 3. Notes: ... CH
14 Chapter Test . CH
14 Acid Base Basics .
CH 14 HC Quiz .

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New Page 1
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Chapter Test B

Chapter: Atoms: The Building Blocks of Matter PART I On the line at the left of each statement, write the letter of the choice that best completes the statement or best answers the question.

1. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude

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that the rays are
composed of a. energy

**Chapter Test B -
PC\|MAC**

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11.1 The Mysterious
Electron Goals To
explain why it is very
difficult to describe the
modern view of the
electron. To give you
some understanding of

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continued 16 Modern
chemistry chapter 3
test b answers. The
measure of the ability
of an atom in a
chemical compound to
attract electrons from
another atom in the
compound is called

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_____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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Chapter 3 Test B
Answers**

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