

## *Chapter 7 Chemical Formulas And Compounds Test B*







### Chapter 7 Chemical Formulas And

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is  $x$  amu, the molar mass is  $x$  g/mol. False b. Samples of equal numbers of moles of two different chemicals

### 7 Chemical Formulas and Chemical Compounds

Chapter 7 Objectives •Explain the significance of a chemical formula. •Determine the formula of an ionic compound formed between two given ions. •Name an ionic compound given its formula. •Using prefixes, name a binary molecular compound from its formula. •Write the formula of a binary molecular compound given its name. Section 1 Chemical Names and ...

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Chapter Seven [Chemical Formulas and Chemical Compounds] Chapter Eight [Chemical Equations and Reactions] Chapter Nine [Stoichiometry] Chapter Ten [States of Matter] Chapter Eleven [Gases] ... Chapter labs: Determining the empirical formula of magnesium oxide or Percent composition .

### Chapter Seven [Chemical Formulas and Chemical Compounds]

Chapter 7 Section 1 Chemical Names and Formulas Significance of a Chemical Formula, continued •The chemical formula for an ionic compound represents one formula unit—the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). •example: aluminum sulfate —Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>

### Chemical Formulas and Chemical Chapter 7 Compounds

Chapter 7 Chemical Formulas and Names. Use this power point for reference only. It mentions the textbook topics and has a few good examples, but most notes will be covered without a power point. We will not do oxidation numbers at this time. (11/4/11). 7.1 Chemical Names and Formulas....

### PPT - Chapter 7 Chemical Formulas and Names PowerPoint ...

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO<sub>3</sub> a. copper(II) carbonate Na<sub>2</sub>SO<sub>3</sub> b. sodium sulfite (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> c. ammonium phosphate SnS<sub>2</sub> d. tin(IV) sulfide HNO<sub>2</sub> e. nitrous acid 2.

### 7 Chemical Formulas and Chemical Compounds

Modern Chemistry 57 Chemical Bonding CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: \_\_\_\_ a. If the formula mass of one molecule is  $x$  u, the molar mass is  $x$  g/mol. \_\_\_\_ b.

### CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds

\_\_\_\_ 7. A chemical formula includes the symbols of the elements in the compound and subscripts that indicate. a. the number of formula units present. b. the number of atoms or ions of each type. c. the formula mass. d. the charges on the elements or ions. \_\_\_\_ 8. How many atoms of fluorine are in a molecule of carbon tetrafluoride, CF<sub>4</sub>? a. 1 c ...

## chapter 7

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: a. S in  $\text{H}_2\text{SO}_3$  b. S in  $\text{MgSO}_4$  c. S in  $\text{K}_2\text{S}$  d. Cu in  $\text{Cu}_2\text{S}$  e. Cr in  $\text{Na}_2\text{CrO}_4$  f. N in  $\text{HNO}_3$  g. C in  $(\text{HCO}_3)$  h. N in  $(\text{NH}_4)$  2. a.

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### Holt McDougal Modern Chemistry Chapter 7: Chemical ...

Chapter 7: Chemical Formulas and Chemical Compounds Section 7-1: Chemical Names and Formulas 7-1-1 Explain the significance of a chemical formula. The subscript indicates that there are 8 carbon atoms in the molecule of octane. The subscript indicates that there are 18 hydrogen atoms in the molecule of octane. Subscript 2 refers to 2 aluminum ...

### Chapter 7: Chemical Formulas and Chemical Compounds

chemical formula reveals the number of atoms of each element contained in a single molecule of the compound, as shown below for the hydrocarbon octane. (Hydrocarbons are molecular compounds composed solely of carbon and hydrogen.) CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION 7-1 OBJECTIVES Explain the significance of a chemical ...

### CHAPTER 7 Chemical Formulas and Chemical Compounds

Chapter 7 Section 1 Chemical Names and Formulas Significance of a Chemical Formula, continued •The chemical formula for an ionic compound represents one formula unit—the simplest ratio of the compound's positive ions (cations) and its negative ions (anions). • example: aluminum sulfate —  $\text{Al}_2(\text{SO}_4)_3$

### Chemical Formulas and Chemical Chapter 7 Compounds

[PDF]Free Chapter 7 Chemical Formulas And Chemical Compounds download Book Chapter 7 Chemical Formulas And Chemical Compounds.pdf Chapter 9 Practice Test - Naming and Writing Chemical Formulas Wed, 17 Apr 2019 05:04:00 GMT Chapter 9 Practice Test - Naming and Writing Chemical Formulas Matching Match each item with the correct statement below.

### Chapter 7 Chemical Formulas And Chemical Compounds

7. What is the empirical formula for lactic acid, which has a molecular formula of  $\text{C}_3\text{H}_5\text{O}_3$ ? a.  $\text{C}_3\text{H}_5\text{O}_3$  b.  $\text{CHO}$  c.  $\text{C}_3\text{H}_5\text{O}_3$  d.  $\text{CHO}$  8.  $\text{NaCl}$  is a. a molecular formula only. b. an empirical formula only. c. both an empirical formula and a chemical formula. d. both a molecular formula and an empirical formula. 9.

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Chapter 7 - Chemical Formulas & Chemical Compounds Chapter 7 is the final chapter of the first semester. It serves as a summary of all course content discussed so far, as the focus of this chapter is on process skills rather than new concepts or theory.

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